UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
General Certificate of Education Ordinary Level

CANDIDATE NAME	

CENTRE NUMBER						CANDIDATE NUMBER				

COMBINED SCIENCE   **5129/02**

Paper 2   **October/November 2008**

2 hours 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen.
You may use a soft pencil for any diagrams, graphs or rough working.
Do not use staples, paper clips, highlighters, glue or correction fluid.
DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.
A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together.
The number of marks is given in brackets [ ] at the end of each question or part question.

For Examiner's Use

This document consists of **17** printed pages and **3** blank pages.

UNIVERSITY *of* CAMBRIDGE
International Examinations

[Turn over

1   The following is a list of gases.

                ammonia   argon   chlorine   hydrogen

                        nitrogen   sulphur dioxide

Complete the following sentences using the gases from the list.

Each gas may be used once, more than once, or not at all.

(a) A gas that causes erosion of buildings is ....   [1]

(b) A gas that displaces bromine from potassium bromide is ....

                                                                                                                                [1]

(c) A gas used in the manufacture of margarine is ....   [1]

(d) A gas used in the purification of water supplies is ....   [1]

(e) A gas that dissolves in water to give an alkaline solution is ....

                                                                                                                                [1]

2   A student connects a cell, a lamp, a variable resistor and an ammeter in series.

(a) In the space below, draw the circuit diagram.

                                                                                                                                [3]

(b) For one setting of the variable resistor, the lamp has normal brightness. The student
     increases the resistance of the variable resistor.

     State what, if anything, happens to

     (i)   the potential difference across the lamp, ...

     (ii) the brightness of the lamp. ....
                                                                                                                                [2]

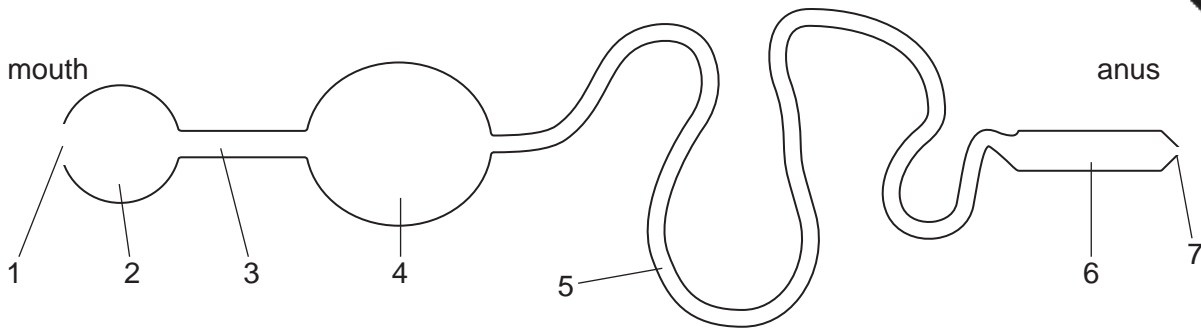**3** Fig. 3.1 represents the human alimentary canal.

mouth   anus

1   2   3   4   5   6   7

Fig. 3.1

(a) State the number of the region where each of the following processes occurs.

(i) absorption ....

(ii) digestion ....

(iii) egestion ....

(iv) ingestion ....   [4]

(b) State

(i) the process that moves food through region 5,

.... [1]

(ii) the part of a balanced diet that helps this process.

.... [1]

(c) Saliva is secreted into the mouth.
State two functions of saliva.

1. ....

2. .... [2]

4 Air is a mixture of gases.

(a) Fig. 4.1 shows the percentages of different gases in air.

0.05% carbon dioxide

0.95% water vapour and other gases

20% gas **A**

79% gas **B**
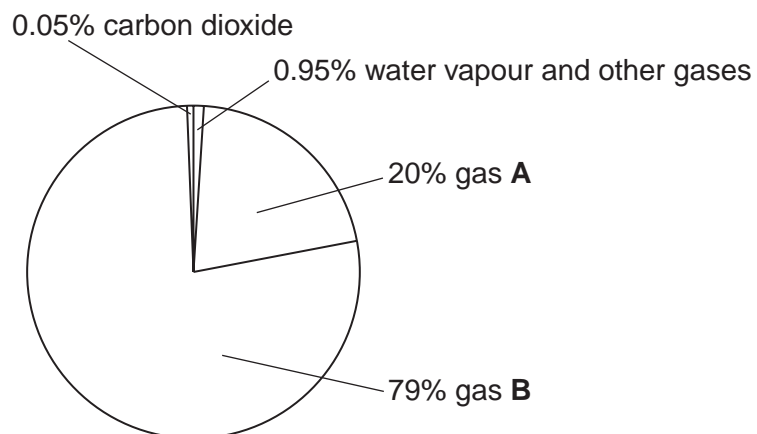
Fig. 4.1

State the name of

gas **A**, ...

gas **B**. ....

[2]

(b) Suggest why the amount of water vapour present in air is **not** constant.

....

.... [1]

(c) Name a pollutant gas that may be found in air.
State the source of this gas.

pollutant gas ....

source ....

[2]

5 An electric lamp produces heat and light.

(a) Name the form of energy that is converted into heat and light. .... [

(b) A lamp is rated as 100W.

(i) Calculate the amount of energy changed into heat and light in the lamp in 5.0 minutes.

amount of energy = ... J [2]

(ii) The lamp produces 28500 J of heat in the 5.0 minutes.

Calculate the amount of energy converted into light by the lamp.

amount of energy = ... J [1]

6 **(a)** Insulin is produced by the pancreas. It acts in the liver, where it affects the conversion of blood glucose to stored carbohydrate.

(i) What type of substance is insulin?

.... [1]

(ii) How is insulin transported to the liver?

.... [1]

(iii) Which organ destroys the insulin after it has had its effect?

.... [1]

(b) Name the part of the eye that changes when

(i) a bright light suddenly shines on the eye,

....

(ii) the eye adjusts from viewing a near object to viewing a distant object.

....
[2]

7 Alkanes are obtained by the fractional distillation of petroleum (crude oil).
Alkenes are made by cracking alkanes.

(a) State **one** condition used in the cracking of alkanes. .... [1]

(b) State how the molecular structure of alkenes differs from the molecular structure of alkanes.

.... [1]

(c) Alkenes may be distinguished from alkanes by adding aqueous bromine.
State the colour change when bromine is added to an alkene.

colour changes from .... to .... [2]

(d) Poly(ethene) is made from ethene by polymerisation.
Explain the meaning of the term *polymerisation*.

....

.... [2]

8 A student performs an experiment to obtain an extension-load graph for a spring.

(a) State what measurements are made in order to determine the extension of the spring.

....

....

.... [1]

(b) The length of the spring with no load is 8.0 cm.
Fig. 8.1 shows the length of the spring for different loads.
Complete Fig. 8.1 by writing in the missing values of extension and load.

load/N	length/cm	extension/cm
1.0	9.2	
	15.2	7.2

Fig. 8.1

[2]

9 (a) The cotyledons of a seed contain an enzyme that digests starch.

   (i) Name an enzyme that digests starch.

   ....

   (ii) Name the product of this digestion.

   ....

                                           [2]

(b) A piece of cotyledon was placed on a jelly containing starch.
A second piece was boiled in water and cooled before being placed on the jelly.

After 24 hours, iodine solution was poured onto the jelly.
Iodine and starch react to give a dark blue/black colour.
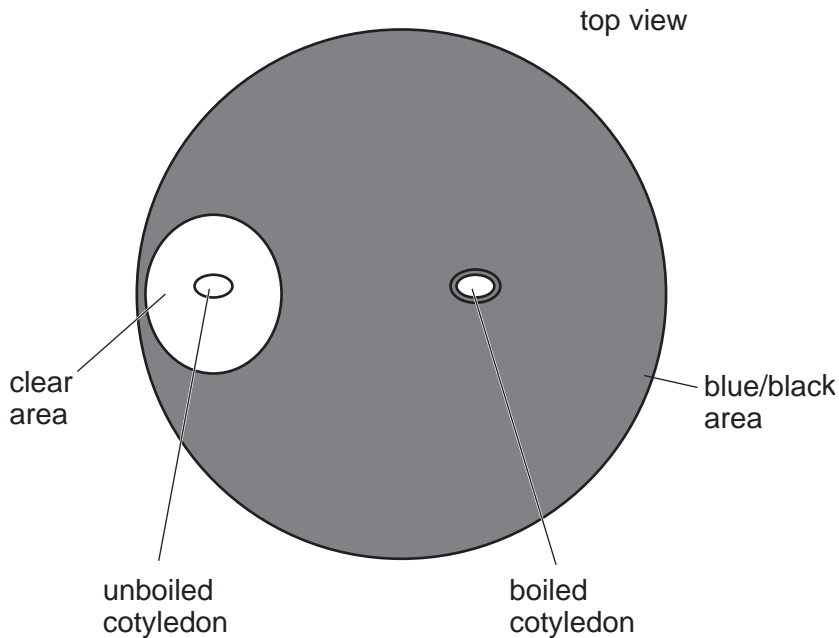Fig. 9.1 shows how the dish looked.

top view

clear
area

blue/black
area

unboiled
cotyledon

boiled
cotyledon

Fig. 9.1

Explain why there is a clear area around the unboiled cotyledon.

....

....

.... [2]

(c) Explain why there is no clear area around the boiled cotyledon.

....

.... [2]

10 Study the following reaction scheme.

```
┌──────────────┐   acid A    ┌──────────────┐      ┌──────────────┐      ┌──────────────┐
│  copper(II)  │───────────▶ │  copper(II)  │  +   │  colourless  │  +   │    water     │
│  carbonate   │             │   sulphate   │      │    gas B     │      │              │
└──────────────┘             └──────────────┘      └──────────────┘      └──────────────┘
```

silver powder   iron powder

```
┌──────────────┐                    ┌─────────────────┐      ┌──────────────────┐
│ no reaction  │                    │ iron(II) sulphate │  +  │  brown solid C   │
└──────────────┘                    └─────────────────┘      └──────────────────┘
```
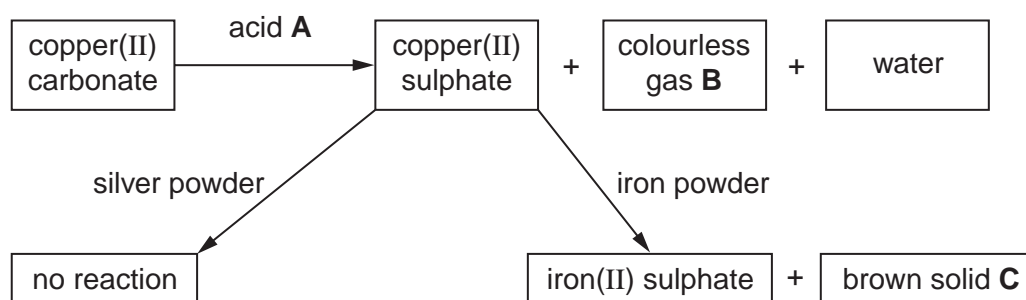
(a) Identify the substances **A**, **B** and **C**.

acid **A** ....

colourless gas **B** ....

brown solid **C** .... [3]

(b) Suggest why the copper(II) sulphate solution reacts with iron but does **not** react with silver.

....

....

.... [2]

11 Fig. 11.1 shows a magnet moving into a coil of wire.
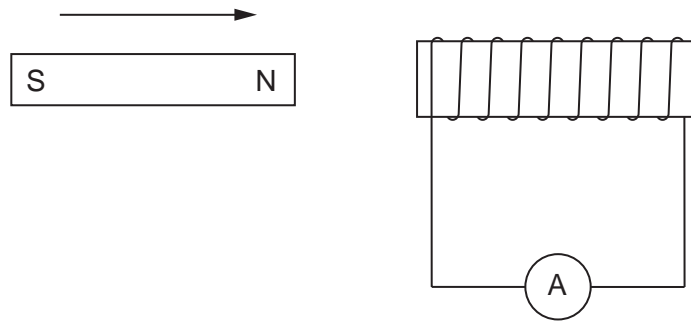   A sensitive ammeter measures the current in the coil.

Fig. 11.1

Complete the following sentences using the words below.

You may use each word once, more than once, or not at all.

     broken   **changing**   **constant**   **complete**   **high**

        induces   **low**   **produces**   **zero**

As the magnet moves into the coil, the .... magnetic field inside the coil

.... an e.m.f. in the circuit. Because the circuit is .... ,

a current is produced. The current is zero whenever the speed of the magnet is

.... [4]

12 Fig. 12.1 represents part of the carbon cycle.
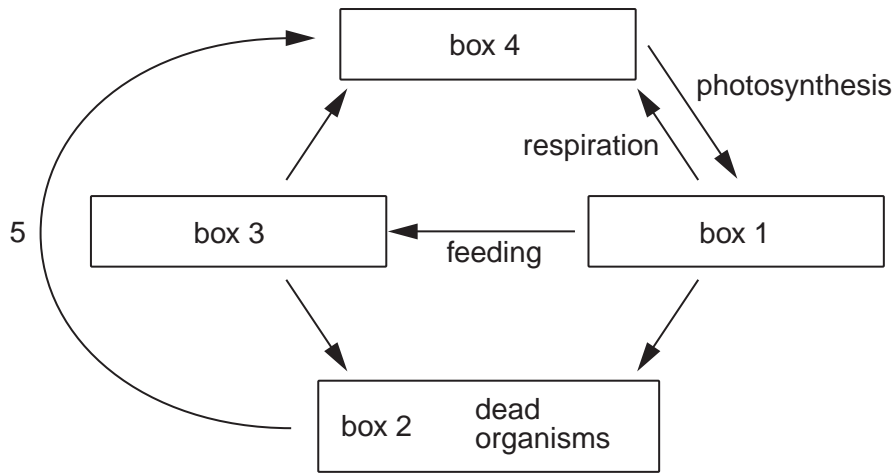
Fig. 12.1

(a) (i) What words should be written in the boxes labelled 1, 3 and 4?

box 1 ....

box 3 ....

box 4 ....
[3]

(ii) State the process that is represented by the arrow labelled 5.

.... [1]

(b) Name the substance in box 1 that contains carbon and is formed by photosynthesis.

.... [1]

13   Fig. 13.1 shows a force of 20 N being used to move a block of mass 5.0 kg across a ho[...] frictionless surface.

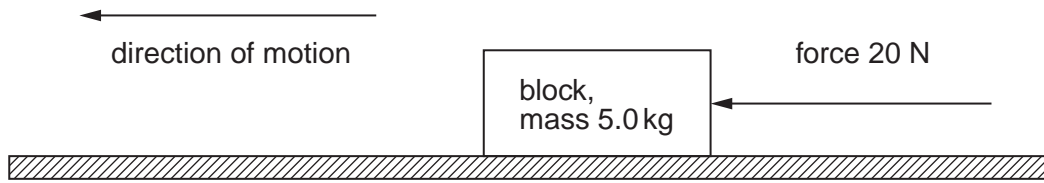direction of motion   force 20 N

block,
mass 5.0 kg

Fig. 13.1

(a) Calculate the acceleration of the block.

acceleration = ... [3]

(b) On Earth, the gravitational field strength $g = 10$ N/kg.

Calculate the weight of the block.

weight = ... N [1]

(c) The force of 20 N moves the block a distance of 40 cm.

Calculate the work done by the force.

work done = ... [2]

14 A rocket explodes 99 m above the ground. Both light and sound are produced at the same time.

The speed of sound is 330 m/s.

(a) (i) Explain why the explosion of the rocket is seen by a boy on the ground before he hears it exploding.

....

.... [1]

(ii) Calculate the shortest time that the sound takes to reach the ground.

time = ... s [2]

(b) The light emitted by the exploding rocket has many different frequencies.

State what is meant by the *frequency* of a wave.

....

.... [1]

(c) Visible light is one component of the electromagnetic spectrum.

Name a component

(i) with a longer wavelength than visible light, ...

(ii) that is emitted from some radioactive nuclei. ....

[2]

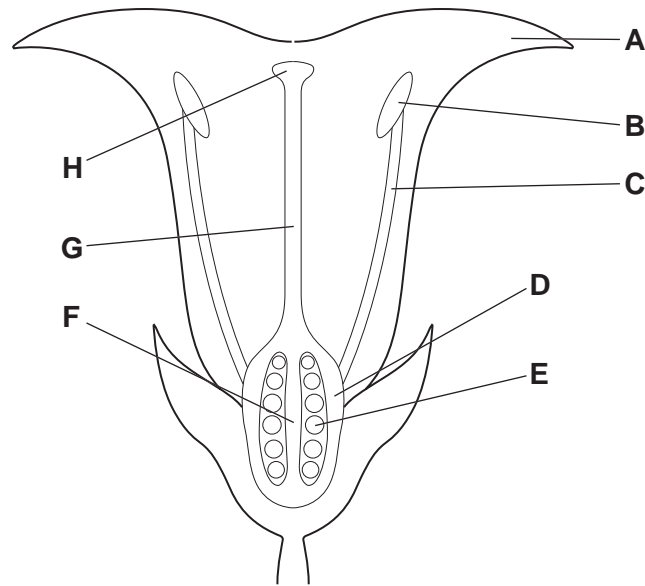**15** Fig. 15.1 shows a cross-section through a flower.

Fig. 15.1

(a) (i) Give the letters of the structures that contain cells with nuclei which fuse during sexual reproduction?

.... and .... [2]

(ii) Name the structure that is formed by the fusion of the nuclei of these two cells.

.... [1]

(b) Into what do each of the structures labelled **D** and **E** develop?

D develops into ....

E develops into ....

[2]

16 ^{16}O and ^{18}O are isotopes of the element oxygen.

(a) State how the two isotopes are different.

....

.... [1]

(b) Explain why the two isotopes have the same chemical properties.

....

.... [2]

(c) State the electronic structure of ^{16}O. .... [1]

(d) State the relationship between the electronic arrangement of an oxygen atom and the group that oxygen belongs to in the Periodic Table.
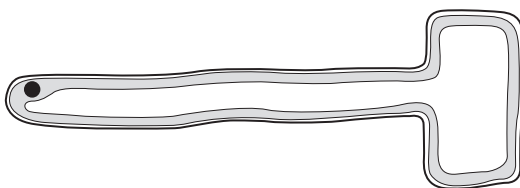
....

.... [1]

17 Fig. 17.1 shows a root hair cell.

Fig. 17.1

(a) State two substances that are absorbed by a root hair cell.

1. ....

2. ....
[2]

(b) State two features of a root hair cell that make it suitable for absorption.

1. ....

....

2. ....

....
[2]

18 **(a)** Fig. 18.1 shows a ray of light from a pin that is incident on a plane mirror at point
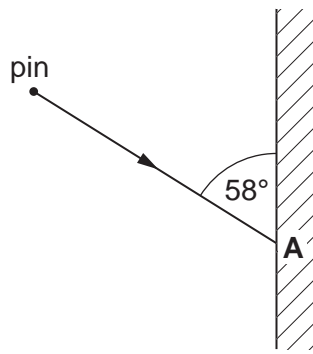
pin

58°

A

Fig. 18.1

   (i) On Fig. 18.1, draw the normal to the mirror at **A**. [1]

  (ii) On Fig. 18.1, mark the position of the image of the pin as seen in the mirror. [2]

 (iii) Calculate the angle of incidence.

angle of incidence = ...° [1]

(b) Fig. 18.2 shows a ray of light entering a parallel-sided glass block.
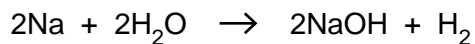
glass block

Fig. 18.2

Complete Fig. 18.2 to show the path of the ray inside the block and after leaving the

19 (a) Define *relative atomic mass*, A_r.

....

.... [2]

(b) Sodium reacts with water to produce sodium hydroxide solution and hydrogen.
The equation for the reaction is

$$2Na + 2H_2O \rightarrow 2NaOH + H_2$$

(i) Calculate the relative molecular mass of sodium hydroxide.
(A_r: H, 1; Na, 23; O, 16.)

.... [1]

(ii) Calculate the mass of water required to react with 4.6 g of sodium.
(A_r: Na, 23.)

mass = ... g [2]

(c) State a test to show that sodium hydroxide solution is alkaline.

test ....

result ....
[2]

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DATA SHEET
The Periodic Table of the Elements

Group

I	II												III	IV	V	VI	VII	0
							1 **H** Hydrogen 1											4 **He** Helium 2
7 **Li** Lithium 4	9 **Be** Beryllium 4												11 **B** Boron 5	12 **C** Carbon 6	14 **N** Nitrogen 7	16 **O** Oxygen 8	19 **F** Fluorine 9	20 **Ne** Neon 10
23 **Na** Sodium 12	24 **Mg** Magnesium 12												27 **Al** Aluminium 13	28 **Si** Silicon 14	31 **P** Phosphorus 15	32 **S** Sulphur 16	35.5 **Cl** Chlorine 17	40 **Ar** Argon 18
39 **K** Potassium	40 **Ca** Calcium 20	45 **Sc** Scandium 21	48 **Ti** Titanium 22	51 **V** Vanadium 23	52 **Cr** Chromium 24	55 **Mn** Manganese 25	56 **Fe** Iron 26	59 **Co** Cobalt 27	59 **Ni** Nickel 28	64 **Cu** Copper 29	65 **Zn** Zinc 30		70 **Ga** Gallium 31	73 **Ge** Germanium 32	75 **As** Arsenic 33	79 **Se** Selenium 34	80 **Br** Bromine 35	84 **Kr** Krypton 36
85 **Rb** Rubidium	88 **Sr** Strontium 38	89 **Y** Yttrium 39	91 **Zr** Zirconium 40	93 **Nb** Niobium 41	96 **Mo** Molybdenum 42	**Tc** Technetium 43	101 **Ru** Ruthenium 44	103 **Rh** Rhodium 45	106 **Pd** Palladium 46	108 **Ag** Silver 47	112 **Cd** Cadmium 48		115 **In** Indium 49	119 **Sn** Tin 50	122 **Sb** Antimony 51	128 **Te** Tellurium 52	127 **I** Iodine 53	131 **Xe** Xenon 54
133 **Cs** Caesium	137 **Ba** Barium 56	139 **La** Lanthanum 57 *	178 **Hf** Hafnium 72	181 **Ta** Tantalum 73	184 **W** Tungsten 74	186 **Re** Rhenium 75	190 **Os** Osmium 76	192 **Ir** Iridium 77	195 **Pt** Platinum 78	197 **Au** Gold 79	201 **Hg** Mercury 80		204 **Tl** Thallium 81	207 **Pb** Lead 82	209 **Bi** Bismuth 83	**Po** Polonium 84	**At** Astatine 85	**Rn** Radon 86
Fr Francium	226 **Ra** Radium 88	227 **Ac** Actinium 89 †																

58-71 Lanthanoid series		140 **Ce** Cerium 58	141 **Pr** Praseodymium 59	144 **Nd** Neodymium 60	**Pm** Promethium 61	150 **Sm** Samarium 62	152 **Eu** Europium 63	157 **Gd** Gadolinium 64	159 **Tb** Terbium 65	162 **Dy** Dysprosium 66	165 **Ho** Holmium 67	167 **Er** Erbium 68	169 **Tm** Thulium 69	173 **Yb** Ytterbium 70	175 **Lu** Lutetium 71
90-103 Actinoid series		232 **Th** Thorium 90	**Pa** Protactinium 91	238 **U** Uranium 92	**Np** Neptunium 93	**Pu** Plutonium 94	**Am** Americium 95	**Cm** Curium 96	**Bk** Berkelium 97	**Cf** Californium 98	**Es** Einsteinium 99	**Fm** Fermium 100	**Md** Mendelevium 101	**No** Nobelium 102	**Lr** Lawrencium 103

Key

a
X
b

a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

The volume of one mole of any gas is 24 dm^3 at room temperature and pressure (r.t.p.).